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(54) Title: METHOD FOR PREPARATION OF CALCIUM-BINDING POLYCARBOXY COMPOUNDS BASED ON POLYSACCHARIDES, AND REPLACEMENTS FOR PHOSPHATES IN DETERGENTS, BASED ON THESE POLYCARBOXY COMPOUNDS

(57) Abstract

The invention provides a method for the preparation of a calcium-complexing polycarboxy compounds by oxidising inulin in the presence of a low concentration of hypohalite, resulting in the production of a polycarboxyinulin containing about 1.2-2.6 carboxyl groups per monosaccharide unit. The hypohalite is preferably hypobromite, which can be produced in situ by chemical or electrochemical oxidation of a catalytic amount of bromide. The polycarboxyinulin can be used as a phosphate substitute in calcium-binding agents and detergents.

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Method for the preparation of calcium-binding polycarboxy compounds based on polysaccharides, and replacements for phosphates in detergents, based on these polycarboxy compounds.

5 The invention relates to a method for the preparation of a calcium-complexing polycarboxy compound based on a polysaccharide.

10 From a technical standpoint, phosphates are the most suitable as calcium-binding agents ("builders") in detergents. However, because of the eutrophication of surface waters, the use of phosphates is increasingly less desirable and there is therefore a need for agents to replace phosphates. Such agents must be able to bind calcium, and preferably also magnesium, well and, in addition, be biodegradable, non-toxic and readily soluble in water and have no adverse influence on the action of the detergent and must also be inexpensive. Many
15 agents have already been investigated, but only a few, such as nitrilotriacetate, zeolites and polycarboxylates, meet the above-mentioned requirements to an extent which renders them also usable in practice. However, doubts are still being expressed as to the safety of nitrilotriacetate. Zeolites, polycarboxylates and nitrilotriacetate
20 are not degradable, or are degradable to an inadequate extent, and, moreover, zeolites have the disadvantage that, because of the necessary "carrier" effect, they form usable "builders" only in combination with other agents. Polyacrylate is currently commonly used as a "co-builder" of this type.

25 Oxidised carbohydrates have also been proposed as substitutes for phosphates. Polydicarboxysaccharides in particular have suitable characteristics for this purpose. These compounds can be obtained by oxidation of carbohydrates, for example with periodic acid followed by chlorite, or directly with hydrogen peroxide or sodium hypochlorite.
30 During this oxidation, the two vicinal hydroxyl functions in the cyclic hexose unit (for example glucose) of the polysaccharide are converted into two carboxyl functions, with ring opening. Methods of this type have been disclosed, for example, in Netherlands Patent Application 70.12380, and M. Diamantoglou et al., Tenside Detergents
35 14, 250-256 (1977) and M.S. Nieuwenhuizen et al., Starch/Stärke 37, 192-200 (1985).

 The known methods for the oxidation of polysaccharides have a number of disadvantages. The use of periodic acid is fairly expensive because of the necessary regeneration of the periodic acid and is also

laborious because the polysaccharide dialdehyde formed with periodic acid as an intermediate has to be further oxidised to the dicarboxylic acid using other agents, such as chlorite or hypochlorite. Hydrogen peroxide is also a relatively expensive oxidising agent and, moreover, makes the use of a catalyst, such as a tungsten compound, desirable for a sufficiently effective conversion; in addition, the low pH required for this reaction is a drawback because of the depolymerisation which occurs. The reaction with hydrogen peroxide is also not very specific, so that the 6-hydroxymethyl group is also oxidised to some extent and significant depolymerisation (chain shortening) takes place.

Polysaccharides that are oxidized with the known methods are always based on glucose, such as 1,4- α -glucans or -polyanhydroglucoses (starch, amylose, amylopectin and substances derived therefrom) and 1,4- β -glucans (cellulose and derivatives thereof). Only Nieuwenhuizen et al. (Starch/Stärke 37, 192-200 (1985)) also investigated the oxidation of other polysaccharides, namely inulin (polyanhydrofructose) and alginate (polyanhydromannuronate/guluronate) with periodate/chlorite; the oxidation products thus obtained from these polysaccharides were, however, found to have poorer calcium-complexing properties than the oxidation products of polysaccharides based on glucose, such as starch and starch derivatives.

It has now been found that inulin can be oxidised in such a way that a polycarboxysaccharide is obtained which has a surprisingly high calcium- and probably magnesium- binding power, combined with a good biodegradability.

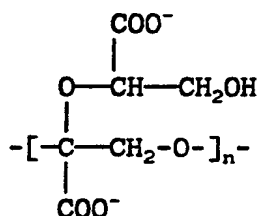
The method according to the invention is therefore characterised in that a polycarboxyinulin is prepared by oxidation of inulin in the presence of a low concentration of hypohalite.

Inulin is a polysaccharide which mainly consists of fructose units. Naturally occurring inulin is a 2,1- β -fructan with a terminal glucose unit. The structure of inulin is shown in the figure. The degree of polymerisation (DP) is highly dependent on the source of the inulin and can range from 6 to 100 or more. Inulin occurs, inter alia, in Jerusalem artichokes (*Helianthus tuberosus*), artichokes (*Cynara scolymus*), chicory (*Cichorium intybus*), dahlias (*Dahlia sp.*) and dandelions (*Taraxacum officinale*). Inulin from Jerusalem artichokes has an average degree of polymerisation of about 15-30, depending on the stage of growth of the plant. The structure and possible uses of

inulin are described, for example, by A. Fuchs and A.G.J. Voragen, Koolhydraten in Nederland (Carbohydrates in the Netherlands), 1988, 13-18.

Where inulin is mentioned in this description, this signifies any polysaccharide which to a substantial extent consists of 2,1- β -anhydrofructose units. Thus, the invention also applies to sinistrin, a branched inulin.

With the method according to the invention, the polycarboxyinulin which, depending on the reaction conditions chosen, contains on average about 1.5 to 2 carboxyl groups per fructose unit is obtained in virtually quantitative yield. This polycarboxyinulin has a degree of polymerisation which is little lower than that of the inulin used as starting material and is readily soluble in water and has good biodegradability. The structure of a polycarboxyinulin having 2 carboxyl groups per fructose unit can be represented schematically by the formula:



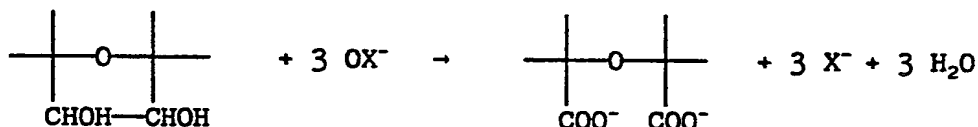
The calcium-binding power of the polycarboxy-inulin prepared in this way is better than that of the polycarboxyamyllose and polycarboxymaltodextrin obtained in accordance with the known method, which latter compounds, in turn, belong to the best calcium-binding agents amongst the modified carbohydrates to date. The product obtained using the method under consideration is therefore outstandingly suitable as a replacement for phosphate in detergents, or as a cobuilder together with, for example, zeolite NaA, and for other similar applications.

The oxidation of polysaccharides with hypohalite, in particular hypochlorite, is known per se. Hypochlorite is a relatively inexpensive oxidising agent and leads to the desired dicarboxy product in one step. However, with the known methods the yield from the oxidation with hypochlorite is fairly low and incomplete oxidation and, additionally, depolymerisation occur and the characteristics of the product are unsatisfactory. Furthermore, in the known conversions, an approximately three-fold excess of hypochlorite is used, which is a disadvantage economically and ecologically.

It has been found that the said disadvantages do not occur or occur to a much smaller extent if a low concentration of oxidising agent is maintained during the oxidation reaction. A further improvement is achieved if hypobromite and/or hypoiodite is used as the actual oxidising agent. In contrast to the known oxidations of polysaccharides with hypochlorite, a stoichiometric amount of oxidising agent can then suffice. These measures are described in the earlier Dutch patent application 8902786, which was not published before the filing date of the present application. Furthermore, it is found that the oxidation proceeds more favourably because inulin is readily soluble in water.

The amount of oxidising agent required is appreciably smaller than in the case of the methods according to the prior art: according to the invention, less than 1.2 times the theoretical amount is used, compared with about 1.5-3 times this amount in accordance with the prior art.

In principle, three molecules of hypohalite are required per fructose unit for the oxidation of C₃-C₄ hydroxyl groups of inulin, in accordance with the empirical equation (in which X represents halogen):



A more extensive oxidation of inulin is accompanied by oxidation of the hydroxymethyl group (C₆) of the fructose unit. However, the presence of such more highly oxidised units in the product does not have an adverse influence on the calcium-binding properties of the product. An incomplete oxidation of the C₃-C₄ hydroxyl groups also still provides a usable oxidation product. The inulin is oxidised in such a way that the oxidation product contains on average 1.0-3.0 and preferably 1.2-2.6 carboxyl groups per original fructose unit.

The oxidation of inulin can be carried out using hypohalite, in particular hypobromite or hypoiodite, as such, which is preferably supplied in the form of salt, for example as sodium hypobromite or potassium hypoiodite, but may also in part be supplied in acid form. In the text which follows, for reasons of clarity, mention is made only of hypobromite, bromine, bromide etc., but the same also applies in principle for hypoiodite, iodine, iodide etc.

Preferably, the hypobromite is prepared in the reaction medium, for example by oxidation of a suitable bromine compound or by disproportionation of bromine to bromide and hypobromite. Since, in principle, there is back-formation of bromide on the oxidation of the polysaccharide by hypobromite, a smaller, for example catalytic, amount of bromide in the reaction mixture can then suffice.

The hypobromite can also be prepared in situ by oxidation of a bromide, for example sodium bromide, by electrochemical means, that is to say by withdrawal of electrons. In this case also there is back-formation of bromide during the oxidation and a catalytic amount therefore suffices. This oxidation can be carried out, for example, by passing an electric current for some time through a solution of the inulin bromide and, if appropriate, a base in an electrochemical cell which is known per se. Preferably, inert electrodes, for example made of graphite or platinum alloys or platinum on titanium, are used in this cell. By means of a suitable choice of electrical voltage and/or electrolyte concentration it can be ensured that the concentration of the oxidising agent remains low. The advantage of electrochemical oxidation is that the product contains much less salt ballast.

Furthermore, it has been found that hypobromite can be produced in a suitable concentration in the reaction medium by oxidation of bromide with chlorine-containing oxidising agents, such as chlorine, hypochlorite, chlorite, etc., in particular hypochlorite. With this method, the concentration of oxidising agent is kept low, for example by adding the hypochlorite solution slowly dropwise to the reaction mixture or by passing chlorine gas slowly through the reaction mixture.

A catalytic amount of bromide is understood to be an amount less than, for example, 20% of bromide in equivalents relative to the oxidation equivalents required. A larger amount is not harmful, but has no additional advantages worthy of note. Preferably, 0.5-10 equivalents of bromide are used per 100 theoretically required oxidation equivalents.

The bromide used can in principle be any bromide-containing salt. For reasons of cost, sodium bromide is preferred. In the case of the oxidation with hypochlorite, the bromide can preferably be added to the inulin but may optionally also be added to the hypochlorite.

An increased reaction rate can be obtained by adding a catalytic amount of a salt of cobalt, manganese, copper, nickel or iron (ca.

0.01 - 1 %). This rat increasing effect was already reported for the oxidation of cellulose (T.P. Nevell and O.P. Singh, Text. Res. J. 56, 270 (1986); O.P. Singh, Text. Dyer Printer 15, 35 (1982); V.A. Shenai and R.B. Prasad, Text. Res. J. 42, 603 (1972) and 44, 671 (1974)) and the oxidation of starch (J. Potze and P. Hiemstra, Starch/Stärke 15, 217 (1963)).

The oxidation can be carried out in a conventional manner, i.e. for example at a temperature of between 10 and 25°C, a pH of 7-11 and with slow addition of the oxidising agent, such as hypochlorit (usually dissolved in water), to the inulin.

If necessary, the temperature can be raised during the reaction in order to achieve a suitable reaction time and degree of conversion. Surprisingly, the result of the oxidation reaction is found to be better if during at least part of the reaction time a temperature of at least 35°C, and preferably of at least 40°C, is maintained. In this context it is possible, for example, to allow the reaction to start at room temperature and to raise the temperature after some time, for example after 1 to 4 hours. It is also possible, with a suitable rat of addition of the oxidising agent, to use such an elevated temperature during the entire reaction.

The reaction mixture is usually worked up by concentration and separation, for example by adding methanol and filtering off the precipitated product. A low pH of, for example, 6-8 can be maintained during working up. Working up at a higher pH has the disadvantage that undesired condensation reactions and associated yellow coloration can occur, while a low pH can give rise to hydrolysis and the depolymerisation associated therewith.

The salts, in particular sodium chloride, can also be separated off simply by direct evaporation of the reaction mixture, owing to the good solubility of the oxidised inulin in water at a pH of 6 or higher. Alternatively, the salts can be separated off by dialysis or ultrafiltration. After the salts have been removed in this way, the product is obtained by evaporation and, if necessary, further purification.

The reaction mixture can, however, also be concentrated in its entirety, followed by freeze-drying, spray-drying or a similar procedure. The product then still contains salts, such as sodium chloride and sodium bromide, but this does not have an adverse influence on the calcium-binding properties, so that a virtually equi-

valent calcium-binding agent can be obtained at lower production costs in this way. The working-up of the reaction mixture can even be dispensed with completely, in which case a readily usable, liquid calcium-binding agent is obtained, for example for use in liquid detergents.

The calcium-binding power of polymer substances can be expressed with the aid of the intrinsic complex constant K_{int} , where $K = Z/[Ca](1-Z) = K_{int}e^{cZ}$. In this equation Z is the degree of occupancy of the polymer and c is a constant for the polymer. Conversion of this equation gives $K_{int} = pCa + \log Z/(1-Z) - aZ$, where $a = c \cdot \log e$ (see M. Floor, A.P.G. Kieboom and H. van Bekkum, *Starch/Stärke* 41, 348 (1989)). For the known polydicarboxysaccharides obtained by oxidation, this constant $\log K_{int}$ is at best in the range of 6-7. The products obtained using the method under consideration have a $\log K_{int}$ of between 7 and 9 or even higher. A parameter which is meaningful in practice for the calcium-complexing characteristics is the calcium-sequestering power (SC), which is defined as the number of mmoles of Ca(II) that can be added to 1 gram of complexing agent until the concentration of non-complex calcium reaches 10^{-5} M (M. Floor, A.P.G. Kieboom and H. van Bekkum, *Starch/Stärke* 41, 348 (1989)). The SC is about 1-1.75 for starch derivatives oxidised with hypochlorite and about 2.5 for starch derivatives oxidised with periodate and chlorite. The products obtained in accordance with the method under consideration have a SC of about 2, i.e. clearly higher than the known oxidised polysaccharides which are economically comparable.

The invention also relates to calcium- and/or magnesium-binding agents which contain polycarboxyinulin obtained in the manner described above. To further improve the calcium-binding properties, these agents can also contain other known "co-builders", such as phosphates, nitrilotriacetate and corresponding compounds such as ethylenediaminetetraacetate and oxydiacetate, and zeolites. In particular, the combination with zeolites yields an outstanding builder for detergents. In a mixture of a polycarboxyinulin with a zeolite, a ratio of 10 : 1 to 1 : 10 can, for example, be used, depending on the intended application, etc.

The invention also relates to detergents which contain the above-mentioned calcium- and/or magnesium-binding agents, or the polycarboxyinulin obtained in the manner described above, in an effective amount.

Example I

1.0 g of inulin (chicory) and 100 mg of sodium bromide were dissolved in 30 ml of water. A sodium hypochlorite solution (10 g Cl_2 /100 ml) was added to the solution in small portions (about 0.5-1 ml).
5 After an immediate rise in pH to about 11, the pH began to fall again. The pH was maintained at a value of 9-9.5 by adding 1 M NaOH. When the fall in the pH became slow, which indicates too low a hypochlorite concentration, fresh hypochlorite solution was added. In total 700 mg of Cl_2 were consumed. The oxidation reaction time was 3 hours. Th
10 product was isolated by concentrating the solution to a volume of 15 ml and pouring into 70 ml of methanol, with stirring. The precipitate was successively filtered off, washed with methanol and dried under reduced pressure. The isolated solution was 1.0 g.

The product has good calcium-binding properties, which can be
15 seen from Table A, in which the most important reaction parameters are also indicated. The structure of the dicarboxy product was confirmed on the basis of the ^{13}C NMR spectrum.

Example II

2.0 g of inulin (chicory), dissolved in 20 ml of water, were
20 oxidised in accordance with the method of Example I, in the presence of 500 mg of NaBr. In total 15 ml of hypochlorite solution were consumed. The oxidation time was 20 hours. The isolated yield was 1.2 g. Both the dissolved product and the isolated product are found to have good calcium-binding properties. See Table A.

Example III

2 g of sinistrin (squill) were oxidised in accordance with the
method of Example I in the presence of 450 mg of NaBr with 16 ml of hypochlorite solution. The oxidation time was 20 hours. The results of the measurement of the calcium-binding power and on the isolated
30 product are given in Table A.

Example IV

5.95 g of inulin (Jerusalem artichoke) and 1.0 g of sodium
bromide were dissolved in 100 ml of water. The inulin was oxidised in the manner described in Example I. 77 ml of hypochlorite solution were
35 added in the course of 2½ hours. After this addition, the solution already displayed good calcium-binding properties; these were found clearly to improve further after a reaction time of 18 hours. The yield of isolated product was 5.0 g. The results are given in Table A.

Example V

Example IV was repeated with the proviso that the reaction temperature was raised to 40-45°C when approximately half of the theoretically required amount of hypochlorite solution had been added. The total reaction time was shortened to 6 hours as a result. The consumption of hypochlorite at the end of the reaction was more than 90% of the theoretical amount. The consumption of alkali was also high (see Table A). The isolated product had an intrinsic complex constant of 9.1.

Example VI

2.0 g of inulin (chicory) were oxidised in accordance with the method of Example I, but without bromide. The hypochlorite consumption was 12 ml instead of 7 ml. The results are given in Table A.

Table A

Example	Temp °C	Time (hours)	eq.HOCl %	Alkali consumption % 1)	Yield % 2)	K _{int}	SC
I	22	3	63	n.d.	50	7.1	1.0
II	22	20	57	45	65	7.8	1.75
III	22	20	61	50	65	9.0	2.0
IV	22	20	98	50	65	7.8	1.75
V	22-45*	6	96	85	90	9.1	2.0
VI	22	3	91	n.d.	n.d.	<4	n.d.

1) Alkali consumption relative to theory, in %

2) Based on a complete conversion with M = 236 of the disodium salt

* After 3 hours at 22° a further 3 hours at 40-45°C

n.d. = not determined

CLAIMS

5 1. Method for the preparation of a calcium-complexing polycarboxy compound based on a polysaccharide, characterised in that a polycarboxyinulin is prepared by oxidation of inulin in the presence of a low concentration of hypohalite.

2. Method according to Claim 1, characterised in that inulin is oxidised in the presence of hypobromite.

10 3. Method according to Claim 2, characterised in that hypobromite is generated in the reaction medium by oxidation of bromide present in a catalytic amount.

4. Method according to Claim 3, characterised in that the bromide is oxidised with hypochlorite.

15 5. Method according to Claim 3 or 4, characterised in that a temperature of at least 35°C is maintained for at least part of the reaction time.

6. Method according to Claim 3, characterised in that the bromide is oxidised electrochemically.

20 7. Method according to one of Claims 1-5, characterised in that inulin is oxidised to a polycarboxyinulin having 1.2-2.6 carboxyl groups per monosaccharide unit.

8. Polycarboxyinulin prepared using the method according to one of Claims 1-7.

25 9. Calcium-binding agent which contains a polycarboxyinulin according to Claim 8.

10. Detergent which contains a polycarboxyinulin according to Claim 8.

INTERNATIONAL SEARCH REPORT

International Application No

PCT/NL 91/00068

I. CLASSIFICATION OF SUBJECT MATTER (if several classification symbols apply, indicate all) ⁶ According to International Patent Classification (IPC) or to both National Classification and IPC IPC ⁵ : C 08 B 37/00																	
II. FIELDS SEARCHED <div style="text-align: center; margin-top: 10px;">Minimum Documentation Searched ⁷</div> <table style="width: 100%; border: none;"> <tr> <td style="width: 20%; border: none;">Classification System</td> <td style="border: none;">Classification Symbols</td> </tr> <tr> <td style="border: none; padding: 10px;">IPC⁵</td> <td style="border: none; padding: 10px;">C 08 B</td> </tr> </table> <div style="text-align: center; margin-top: 10px;">Documentation Searched other than Minimum Documentation to the Extent that such Documents are Included in the Fields Searched ⁸</div>			Classification System	Classification Symbols	IPC ⁵	C 08 B											
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III. DOCUMENTS CONSIDERED TO BE RELEVANT ⁹ <table border="1" style="width: 100%; border-collapse: collapse;"> <thead> <tr> <th style="width: 10%; padding: 5px;">Category ⁹</th> <th style="width: 70%; padding: 5px;">Citation of Document, ¹¹ with indication, where appropriate, of the relevant passages ¹²</th> <th style="width: 20%; padding: 5px;">Relevant to Claim No. ¹³</th> </tr> </thead> <tbody> <tr> <td style="text-align: center; vertical-align: top; padding: 5px;">A</td> <td style="padding: 5px;">J. Agn. Chem. for Japan, volume 28, 1954, K. Maekawa et al.: "Oxidation of inulin with periodate", pages 357-363 see page 363; Summary --</td> <td></td> </tr> <tr> <td style="text-align: center; vertical-align: top; padding: 5px;">A</td> <td style="padding: 5px;">GB, A, 1330123 (UNILEVER) 12 September 1973 cited in the application --</td> <td></td> </tr> <tr> <td style="text-align: center; vertical-align: top; padding: 5px;">A</td> <td style="padding: 5px;">Die Stärke, volume 15, no. 6, J. Potze et al.: "Über den Einfluss der Reaktionsbedingungen auf die Oxydation der kartoffelstärke mit Hypochlorit", pages 217-225, 236 --</td> <td></td> </tr> <tr> <td style="text-align: center; vertical-align: top; padding: 5px;">E</td> <td style="padding: 5px;">EP, A, 0427349 (TNO) 15 May 1991 see the whole document -- ./.</td> <td style="vertical-align: top; padding: 5px;">1-10</td> </tr> </tbody> </table>			Category ⁹	Citation of Document, ¹¹ with indication, where appropriate, of the relevant passages ¹²	Relevant to Claim No. ¹³	A	J. Agn. Chem. for Japan, volume 28, 1954, K. Maekawa et al.: "Oxidation of inulin with periodate", pages 357-363 see page 363; Summary --		A	GB, A, 1330123 (UNILEVER) 12 September 1973 cited in the application --		A	Die Stärke, volume 15, no. 6, J. Potze et al.: "Über den Einfluss der Reaktionsbedingungen auf die Oxydation der kartoffelstärke mit Hypochlorit", pages 217-225, 236 --		E	EP, A, 0427349 (TNO) 15 May 1991 see the whole document -- ./.	1-10
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E	EP, A, 0427349 (TNO) 15 May 1991 see the whole document -- ./.	1-10															
<div style="display: flex; justify-content: space-between;"> <div style="width: 45%;"> <p>¹⁰ Special categories of cited documents: ¹⁰</p> <p>"A" document defining the general state of the art which is not considered to be of particular relevance</p> <p>"E" earlier document but published on or after the international filing date</p> <p>"L" document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)</p> <p>"O" document referring to an oral disclosure, use, exhibition or other means</p> <p>"P" document published prior to the international filing date but later than the priority date claimed</p> </div> <div style="width: 45%;"> <p>"T" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention</p> <p>"X" document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step</p> <p>"Y" document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art</p> <p>"A" document member of the same patent family</p> </div> </div>																	
IV. CERTIFICATION <table border="1" style="width: 100%; border-collapse: collapse;"> <tr> <td style="width: 50%; padding: 5px;"> Date of the Actual Completion of the International Search <div style="text-align: center; margin-top: 10px;">25th July 1991</div> </td> <td style="width: 50%; padding: 5px;"> Date of Mailing of this International Search Report <div style="text-align: center; margin-top: 10px;">30. 08. 91</div> </td> </tr> <tr> <td style="width: 50%; padding: 5px;"> International Searching Authority <div style="text-align: center; margin-top: 10px;">EUROPEAN PATENT OFFICE</div> </td> <td style="width: 50%; padding: 5px;"> Signature of Authorized Officer <div style="text-align: center; margin-top: 10px;"> <div style="border: 1px solid black; display: inline-block; padding: 2px 10px;">M. PEIS</div> M. Pez </div> </td> </tr> </table>			Date of the Actual Completion of the International Search <div style="text-align: center; margin-top: 10px;">25th July 1991</div>	Date of Mailing of this International Search Report <div style="text-align: center; margin-top: 10px;">30. 08. 91</div>	International Searching Authority <div style="text-align: center; margin-top: 10px;">EUROPEAN PATENT OFFICE</div>	Signature of Authorized Officer <div style="text-align: center; margin-top: 10px;"> <div style="border: 1px solid black; display: inline-block; padding: 2px 10px;">M. PEIS</div> M. Pez </div>											
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III. DOCUMENTS CONSIDERED TO BE RELEVANT (CONTINUED FROM THE SECOND SHEET)		
Category *	Citation of Document, ** with indication, where appropriate, of the relevant passages	Relevant to Claim No.
A	<p>Chemical Abstracts, volume 104, no. 20, May 1986, (Columbus, Ohio, US), see page 107, abstract no- 170459g & JP, A, 60226502 (MITSUBISHI GAS CHEMICAL CO., LTD), 11 November 1985</p> <p>-----</p>	

**ANNEX TO THE INTERNATIONAL SEARCH REPORT
ON INTERNATIONAL PATENT APPLICATION NO.**

NL 9100068
SA 46769

This annex lists the patent family members relating to the patent documents cited in the above-mentioned international search report. The members are as contained in the European Patent Office EDP file on 21/08/91
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Patent document cited in search report	Publication date	Patent family member(s)	Publication date
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		BE-A- 755050	19-02-71
		CA-A- 971561	22-07-75
		CH-A- 547852	11-04-74
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		FR-A- 2063914	09-07-71
		GB-A- 1330121	12-09-73
		GB-A- 1330122	12-09-73
		NL-A- 7012380	23-02-71
		SE-A, B 368225	
EP-A- 0427349	15-05-91	NL-A- 8902786	03-06-91

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